Edexcel iGCSE Chemistry Checklist

Section B13: Electrolysis

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| ***I can*** | ☺ | 😐 | ☹ |
| understand that an electric current is a flow of electrons or ions |  |  |  |
| understand why covalent compounds do not conduct electricity |  |  |  |
| understand why ionic compounds conduct electricity only when molten or in solution |  |  |  |
| understand why ionic compounds conduct electricity only when molten or in solution |  |  |  |
| understand that electrolysis involves the formation of new substances when ionic compounds conduct electricity |  |  |  |
| describe experiments to investigate electrolysis, using inert electrodes, of molten salts such as lead(II) bromide and predict the products |  |  |  |
| **describe experiments to investigate electrolysis, using inert electrodes, of aqueous solutions such as sodium chloride, copper(II) sulfate and dilute sulfuric acid and predict the products** |  |  |  |
| write ionic half-equations representing the reactions at the electrodes during electrolysis |  |  |  |
| **recall that one faraday represents one mole of electrons** |  |  |  |
| **calculate the amounts of the products of the electrolysis of molten salts and aqueous solutions.** |  |  |  |
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